Re: chemistry equilibrium

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From: Chris Wiegert (*cwiegert_at_telus.net*)
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> 2NO2(g) <=> N2O4(g)
>
> 138g of N2O4 were put into a vessel of volume 4.00 dm^3. At
> equilibrium 17.25g of NO2 were present. Calculate
> a) the no. of moles initially present and the no. of moles of NO2 at
> equilibrium
> b)the no. of moles of each component at equilibrium
> C)the total no. of moles of all components present
> d)the fractions of each
> e)if the total pressure is 10^6Pa the partial pressure of each
> f)Kp for equilibrium
> g) use your data from b) to find the concentration of each component
> in mol dm^(−3)
> h)find Kc for this equilibrium

This group can help you with your homework, but we will not do it for you. Let's start with a. Do you know how to find the molar mass of those two molecules? You'll need a periodic table that shows the molar mass of each atom. That, plus some simple math, should get you down to d at least.