

# Re: WHICH SCIENCE OR SCIENTISTS CAN YOU TRUST?

**Source:** <http://sci.tech-archive.net/Archive/sci.chem/2005-03/0281.html>

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**From:** Dirk Bruere at Neopax ([dirk\\_at\\_neopax.com](mailto:dirk_at_neopax.com))

**Date:** 03/04/05

Date: Fri, 04 Mar 2005 18:14:43 +0000

farooq\_w@hotmail.com wrote:

> *Dirk Bruere at Neopax wrote:*

>

>

>>>>>> *Uranium exists in topsoil to the extent of about 10 tonnes per*

>

> *sq km.*

>

>>>>>>

>>>>>> *It does ? ... where ?*

>>>>>>

>>>>>>

>>>>>> *Pretty much everywhere.*

>>>>

>>>>

>>>>

>>>> *Gosh. Over what depth in that sq km are you talking here ?*

>>>

>>>

>>> *Top one metre.*

>>

>> *IIRC it's about 2ppm of soil, and I think it was per sq mile, not sq*

>

> *km so it's*

>

>> *a bit less than I said above.*

>>

>> *Bricks are 8ppm, and plasterboard 13ppm of Uranium, if >that's of any*

>

> *use.*

>

> *Where are you quoting these values from? 2mg U/kg soil is a large unit*

> *for analytical chemists. This means common topsoil should test positive*

> *for uranium. The upper limit set for the metal by EPA is 30 parts per*

> *billion in drinking water. If top soil is 2ppm = 2000ppb U, we would*

sci.chem: Re: WHICH SCIENCE OR SCIENTISTS CAN YOU TRUST?

> *all die slowly from it.*

>

<http://www.cartage.org.lb/en/themes/Sciences/Chemistry/NuclearChemistry/Transuranium/Uranium.htm>  
<http://tinyurl.com/66qfu>

"One isotope of thorium, Th-232, has a half – life of 14 billion years and has an abundance of 12 parts per million in the Earth's crust. It is a principal constituent of some minerals, notably thorite and monazite (a mixed rare–earth and thorium phosphate). Three isotopes of uranium are found in nature. Their isotopic abundances and half–lives are U–234, 0.006%, 230,000 years; U–235, 0.72%, 696 million years; and U–238, 99.27%, 4.51 billion years. Some laboratory produced isotopes of uranium also have long halfives. The overall abundance of uranium in the Earth's crust is about 4 parts per million,"

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Dirk

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