

## Re: neutralise citric acid

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*Source:* <http://sci.tech--archive.net/Archive/sci.chem/2005-06/msg00201.html>

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- *From:* Bob <bbx107@xxxxxxxxxxx>
  - *Date:* Mon, 06 Jun 2005 18:43:53 -0700
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On 6 Jun 2005 11:25:59 -0700, stuart\_noble@xxxxxxxxxxx wrote:

>>Simple acid-base calculation (for homework ?).  
>Getting a bit old for homework I'm afraid :-)  
>If the answer is 57.129 gms, I'm happy with that.

It's not. So don't get happy just yet.

First, the person who calculated that made an assumption about what you meant, and that was not what you meant. You want to make the solution neutral (pH 7), rather than neutralize the acid (add an amount of base equal in moles to the amount of acid). (Don't feel bad that the terms are confusing. But note that everyone asked you what you meant.)

So you would need somewhat less than what he calculated. But he also noted that you may well use NaOH that is not pure. So you would need more.

What are you trying to do?? Maybe... add NaOH (carefully!!!), and after each amount has dissolved and you have good mixing, check the pH with pH paper (or meter). The calculation done earlier is a guide, but then add as much as needed to achieve the desired result.

>As for calling my hardware store and asking for trisodium citrate, I  
>think I can envisage the reponse.

?? I thought it was commonly available??

bob

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- *Follow-Ups:*

Re: neutralise citric acid

- ◆ **Re: neutralise citric acid**  
◇ From: stuart\_noble

• **References:**

- ◆ **neutralise citric acid**  
◇ From: stuart\_noble
  - ◆ **Re: neutralise citric acid**  
◇ From: Bevan Price
  - ◆ **Re: neutralise citric acid**  
◇ From: stuart\_noble
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