

Re: molarity

Source: <http://sci.tech-archive.net/Archive/sci.chem/2005-07/msg00563.html>

- *From:* Uncle Al <UncleAl0@xxxxxxxxxxxxxx>
 - *Date:* Tue, 26 Jul 2005 09:10:54 -0700
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John Doe III wrote:

- >
- > Let's assume 100 g of NaCl (MW = 58 g./ mol) is dissolved in 3.5 L of water.
- > Is the molarity
- > A) $100 / (58 * 3.5)$
- > or
- > because NaCl dissociates into 2 ions
- > B) $2 * 100 / (58 * 3.5)$
- >
- > ???
- >
- > What is the molarity of a salt solution?
- >
- > is it the number of moles of all dissociated solutes that are dissolved in
- > the solvent divided by liter of solvent????

You have moles/liter of Cl⁻, Na⁺, and NaCl by formula unit. All three numbers are identical. Cf: molality (correct spelling) and colligative properties.

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Uncle Al

<http://www.mazepath.com/uncleal/>

(Toxic URL! Unsafe for children and most mammals)

<http://www.mazepath.com/uncleal/qz.pdf>

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Re: molarity

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