

Re: molarity

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- *From:* "John Doe III" <john_doeIII@xxxxxxxx>
 - *Date:* Tue, 26 Jul 2005 16:14:40 GMT
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> The problem becomes a bit more tricky when it is a di (or tri) valent
> salt.
> For instance, a 1.0 M solution of sodium carbonate would be 2.0 M in
> sodium
> ions, and 1.0 M in carbonate ions.
> If that makes sense, then you are ahead of most introductory college
> students.

Thanks so much.

So if, in a solution, we have 3 M Na⁺ and 1 M CO₃⁻ is the molarity of this solution:

2 M: 1 M Na₂CO₃ + 1 M Na⁺ that is left = 2M

or

1 M: because Na₂CO₃ is the only thing I should considered

?????

• *Follow-Ups:*

- ◆ **[Re: molarity](#)**
◇ *From:* Craig

• *References:*

- ◆ **[molarity](#)**
◇ *From:* John Doe III
- ◆ **[Re: molarity](#)**
◇ *From:* Dr_Dickie

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