

# Re: molarity

---

*Source:* <http://sci.tech-archive.net/Archive/sci.chem/2005-07/msg00573.html>

---

- *From:* Borek <[borek@xxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxx](mailto:borek@xxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxx)>
  - *Date:* Tue, 26 Jul 2005 20:43:15 +0200
- 

On Tue, 26 Jul 2005 19:47:29 +0200, Craig <[cagerken@xxxxxxx](mailto:cagerken@xxxxxxx)> wrote:

So if, in a solution, we have 3 M Na<sup>+</sup> and 1 M CO<sub>3</sub><sup>-</sup> is the molarity of this solution:

2 M: 1 M Na<sub>2</sub>CO<sub>3</sub> + 1 M Na<sup>+</sup> that is left = 2M  
or  
1 M: because Na<sub>2</sub>CO<sub>3</sub> is the only thing I should considered

Neither, really. You seem to have an idea that "molarity" is a property that can be applied to a solution as a whole. If you had a complicated mixture of many solutes, like, say, seawater or blood, you would rarely say "this is a 2.0 M solution", as if that encompassed everything. There are times when it is useful to know the total concentration of every blessed thing in solution. Most often, we speak of the concentration of a specific component.

To add to that: JD, take a look at the bottle of mineral water, they have sometimes concentrations of dissolved ions given (in mg/L, but it doesn't matter) - every ion concentration is given separately, as if it doesn't matter what was dissolved to obtain the final effect.

Best,  
Borek  
--

<http://www.chembuddy.com> - chemical calculators for labs and education  
BATE - program for pH calculations  
CASC - Concentration and Solution Calculator  
pH lectures - guide to hand pH calculation with examples  
.