

Re: Physical Interpretation of Negative pH.

Source: <http://sci.tech-archive.net/Archive/sci.chem/2005-08/msg00898.html>

- *From:* "Farooq W" <farooq.w@xxxxxxxx>
 - *Date:* 29 Aug 2005 10:08:59 -0700
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Borek wrote:

> On Mon, 29 Aug 2005 17:09:43 +0200, Farooq W <farooq.w@xxxxxxxx> wrote:
>
>> (iii) A short note in J.Chem.Ed has a procedure for making a pH
>> -47 to +47 HCl solution which I think is meant to be a joke.
>>
>> Suppose we have 1 liter of 1×10^{-6} M HCl, add about 99 L of water, the
>> [HCl] is 1×10^{-8} M and pH= 8. Now suppose you have single molecule of
>> HCl in 10^{23} L of water, so that
>> [HCl] = 1.7×10^{-47} , the pH is then 47!
>
> I agree with Marvin – that's not even funny.

Fortunately this is not my view too. The Journal of Chemical Education must be held responsible for spreading such ideas.

>
> <http://www.chembuddy.com/?left=pH-calculation&right=pH-strong-acid-base>
>
> Best,
> Borek
> --
> <http://www.chembuddy.com> – chemical calculators for labs and education
> BATE – program for pH calculations
> CASC – Concentration and Solution Calculator
> pH lectures – guide to hand pH calculation with examples

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- *Follow-Ups:*
 - ◆ ***Re: Physical Interpretation of Negative pH.***
◇ *From:* Borek
 - *References:*
 - ◆ ***Physical Interpretation of Negative pH.***
◇ *From:* Farooq W

Re: Physical Interpretation of Negative pH.

◆ **Re: Physical Interpretation of Negative pH.**

◇ From: Borek

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