

Re: How do I calculate evaporative losses?

Source: <http://sci.tech-archive.net/Archive/sci.chem/2006-11/msg00354.html>

- *From:* "John M." <john_howard_morgan@xxxxxxxxxxxxxx>
 - *Date:* 22 Nov 2006 12:08:08 -0800
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BAJERFAN wrote:

lcc wrote:

On Tue, 21 Nov 2006 22:25:21 +0100, <elephantiasis@xxxxxxxxxxxxxx>
wrote:

I'm looking at a reaction involving methanol for making biodiesel from vegetable oil. Methanol and caustic soda are mixed and added to vegetable oil. The reaction takes place at a range of temperatures between 40 and 65 degrees and usually lasts for an hour. The products are biodiesel and glycerine.

It can take place in a sealed vessel but more usually is done in a vessel with a lid which is not sealed.

I'm expecting some of the methanol will be lost by evaporation to atmosphere but how can I calculate the likely losses?

It's complex matter. Depend on temperature (Methanol boil at 64C), vessel and lid.

Methanol is a poison, it has to be done in ventilated area. Why can't you use condenser ?

Lucas

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Re: How do I calculate evaporative losses?

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Anyone care to speculate on the source of the glycerine since this would be a product of saponification of beef fat but hardly would be expected to arise from hydrolysis of veggie oil. Glycerine is usually found in fats of animal origin. You makin moo-diesel?

All common fats and oils are glycerides. Read:
http://en.wikipedia.org/wiki/Vegetable_oil if you don't believe me