

The difference between a "galvanic cell" and "electrolyte cell"

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Hi

Reading Electrochemistry by Hamann [1998, p.237] he says that If a cell has a non offset (no DC) ac signal applied it will operate like "galvanic cell" rather than an "electrolyte cell".

I was under the impression that they were the same thing. However, a good old search on Google reveled the following difference:

– Galvanic cell: $\Delta G < 0$; spontaneous; $-\Delta G =$ maximum electric work output

– Electrolyte cell: $\Delta G > 0$; non-spontaneous; $\Delta G =$ minimum electric work needed

Does this mean that a Galvanic will produce electrical energy and a electrolyte cell with produce chemical?

Can someone clear this up for me?

Cheers

Wayne