

## Re: Endless Oil?

**Source:** <http://sci.tech-archive.net/Archive/sci.energy/2004-09/1194.html>

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**From:** Dan Bloomquist (EXTRApublish21\_at\_lakeweb.com)

**Date:** 09/29/04

Date: Wed, 29 Sep 2004 02:58:45 GMT

Bill Ward wrote:

- >
- > *The decomposition rate of ozone is only half the equation.*
- > *The production rate is equally important*
- >
- > *My understanding is that ozone (O3) is formed when short UV*
- > *hits O2 in the upper atmosphere. The UV is absorbed in the*
- > *process, thus reducing the amount reaching the surface.*
- >
- > *One thing about photochemical reactions is that they stop*
- > *when the lights go out. During winter at the poles, it's*
- > *continuously dark – no UV, no O3 formation..*
- >
- > *So why is anyone surprised when the [O3] drops during the*
- > *polar winter?*
- >
- > *Can someone fill me in on the production rate and lifetime*
- > *of the O3 layer without Freon? If the O3 layer were*
- > *somehow completely removed, how long would it take to be*
- > *reformed by the solar UV?*

Here is the first faq:

<http://www.faqs.org/faqs/ozone-depletion/intro/>

I think to start with, ozone does not want to be ozone. Leave it alone in a cold dark night and it may survive. Add a catalyst and...

I've got a 'feeling' that it would recover very fast. It is an equilibrium thing. And that is why it is down some 20% ? 30%. That is the equilibrium we now find with the given chlorine radicals up there.

The problem with CFAs is cumulative because the the very long lifetime of these transported radicals.

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- >
- > *Regards,*

Re: Endless Oil?

sci.energy: Re: Endless Oil?

>  
> *Bill Ward*

Best, Dan.

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<http://lakeweb.net>

<http://ReserveAnalyst.com>

No EXTRA stuff for email.