

# Re: Bohr's Atom still number one

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- *From:* "guskz@xxxxxxxxxxxx" <guskz@xxxxxxxxxxxx>
  - *Date:* 1 May 2007 23:00:10 -0700
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On May 1, 5:46 am, bz <bz+...@xxxxxxxxxxxxxxxxxxxxxx> wrote:

"g...@xxxxxxxxxxxx" <g...@xxxxxxxxxxxx> wrote  
[innews:1177995332.605199.251900@xxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxx](mailto:innews:1177995332.605199.251900@xxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxxx):

(b) the s-orbital is spherical and the Bohr orbit is not

Bohr's orbit isn't circular? then what is it squarical?

Circular is not spherical.

A circular orbit has one kind of symetry.

A spherical orbit had a completely different kind of symetry.

An atom with a circular orbit will only bond with atoms that lie on the plane of the orbit.

A atom with a spherical orbit will bond with atoms in any direction.

The chemistry of atoms that have 's' orbitals containing their bonding electron clearly shows spherical symetry, not circular.

Bohr's orbitals ALL lie in the same plane.

Bohr's model is disproved by the bonding of molecules.

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bz

please pardon my infinite ignorance, the set-of-things-I-do-not-know is an infinite set.

bz+...@xxxxxxxxxxxxxxxxxxxxxx remove ch100-5 to avoid spam trap

There is an obediance (non-random and non cloudy/fuzzy) of perfect

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INTEGER symmetry that is "demanded" in the GEOMETRY of the atom for the Table of Elements to obey the quantum principle and for the electron's necessity to shed or gain light at specific intervals in this geometry.

It is a no dice game.

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