

Re: Bohr's Atom still number one

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- *From:* "guskz@xxxxxxxxxxxx" <guskz@xxxxxxxxxxxx>
 - *Date:* 2 May 2007 14:56:30 -0700
-

On May 2, 5:24 am, bz <bz+...@xxxxxxxxxxxxxxxxxxxxxxxx> wrote:

"g...@xxxxxxxxxxxx" <g...@xxxxxxxxxxxx> wrote
innews:1178085610.863662.57900@xx:

On May 1, 5:46 am, bz <bz+...@xxxxxxxxxxxxxxxxxxxxxxxx> wrote:

"g...@xxxxxxxxxxxx" <g...@xxxxxxxxxxxx> wrote
innews:1177995332.605199.251900@xx:

(b) the s-orbital is spherical
and the Bohr orbit is not

Bohr's orbit isn't circular? then what is it
squarical?

Circular is not spherical.

....

Bohr's orbitals ALL lie in the same plane.
Bohr's model is disproved by the bonding of molecules.

There is an obedience (non-random and non cloudy/fuzzy) of perfect

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INTEGER symmetry that is "demanded" in the GEOMETRY of the atom for the Table of Elements to obey the quantum principle and for the electron's necessity to shed or gain light at specific intervals in this geometry.

What 'electron's necessity to shed or gain light at specific intervals'???

Meaning as the electron approaches or moves away from the nucleus a photon is shed or absorbed.

It is a no dice game.

Einstein said something similar. You are no Einstein.

Correct therefore give Einstein some validation, not me.

A sufficiently conformative and simplistic geometry (of forces) is needed since the electron valence is a sufficiently conformative (integer numeral valences) and a simplistic atomic behavior (the opposite of complex atomic behavior patterns).

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bz

please pardon my infinite ignorance, the set-of-things-I-do-not-know is an infinite set.

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