

Re: How can NaCl dissolve in water? Can Any one explain this phenomenon in molecular level?

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- *From:* "newedana" <[simplesong1004@xxxxxxxxxxxxx](mailto:simplesong1004@xxxxxxxxxxxxx)>
  - *Date:* 12 Jun 2005 14:55:36 -0700
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>As I posted on June 12. 3: 27, the force of binding atoms to build a molecules is not the electrostatic Coulomb's attraction believed today, but is the Meissoner's magnetic attraction induced by orbital persistent current, or orbital electron ring. This Meissoner's magnetism has an inherent character, something different from ordinary magnetism. It has a diamagnetic character against the magnetism induced by other electron rings with different orbital radii, because the electron ring hates to be influenced by other magnetism with a different intensity, due to energetic reason. So there acts repulsion between two electron rings with different radii, the diamagnetic phenomenon. If there is no space to repulse each other, they have to arrange to make a right angles between them in order to minimize their mutual magnetic influence. This Meissoner's magnetic behavior governs all the interaction phenomena occurring between materials, without exceptions.

Dissolving phenomenon is also included to be explained with this principle. You posted like dissolves like. That is right! The most intensive attraction has to act between the same kind of atoms that have exactly the same electron rings of all their kinds.

As I posted at the beginning how NaCl can dissolve in water, people replied, are you kidding us? It is quite natural because we have been taught an incorrect science up to the present at the school; molecules are built by electrostatic attraction between atoms by a snapshot effect, and electrolyte materials being divided into several kinds of molecular fragments with counter charges, and etc.

An entirely new view for the dissolution of NaCl in water proposed by Dr. Yoon ([www.yoonsatom.net](http://www.yoonsatom.net)), is that, the non-bonded orbital electron rings in the outermost Cl electron shell attract oxygen of water molecules succesively to build water chains as \*OH2\*H2O\*OH2\*H2O.... Symbol \* signifies the Meissoner's magnetic attraction. Thus NaCl molecule can coexist amid the water molecules. The number of water molecules is determined by NaCl concentration in water. The minimum number of water molecules being associated on Cl atom can be readily estimated from NaCl saturate solution. Under this number of water molecules, even a single NaCl molecule is not allowed to dissolve in water, and begin to be driven out from the water phase. That is why electrolyte materials have a fixed solubility at given temperature and pressure which determines the minimum number of water molecules. Then how can we figure out the shape of Na<sup>+</sup> and Cl<sup>-</sup> ions? Both ions have the same shape, as NaCl(nH2O). So that Cl atom carrying water chains repulse Na being loosely connected to it. This is the so called ion. It is known dry HCl has covalent bond. This covalent bond immediately turns into ionic bond in presence of water molecules. Thus we can conclude that ionic and covalent bonds are essentially not different

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but the same.

Water molecules dissociate from HCl at 108.6 degrees C one atom, the azeotropic boiling temperature. Refer that water boils at 100 degrees C while HCl boils at -84.9 degrees C. We can thus understand also the azeotropic boiling character of materials with the same principle.

CO<sub>2</sub> molecules attach to CO radicals involved in caffeine or nicotine which are non-volatile materials, so these substance can dissolve in CO<sub>2</sub> gas phase to build a solution of gas/solid system. This phenomenon is allegedly utilized to separate these kind of non-volatile macro-molecular substance purely from their mother substance, utilizing CO<sub>2</sub> gas phase under crytical state, however this phenomenon of forming gas/solid solution was impossible to understand with present science.

Mr p6! You said you bought the expensive Dr. Yoon's textbook, but you says as though you do not read his book. I red his book five times for establishing a foundation of understanding his science principles.

People never knows why CaO, CaCO<sub>3</sub>, BaSO<sub>4</sub> and etc. do not readily dissolve in water while NaO, Na<sub>2</sub>CO<sub>3</sub>, Na<sub>2</sub>SO<sub>4</sub>, K<sub>2</sub>SO<sub>4</sub> and etc. dissolves pretty well into water, but I became to understand this reason.

It is not limited to understand dissolution phenomenon of inorganic materials but also all kind of organic substances.

Dr, Yoon explains how raw egg-white can dissolve into water despite that it has a long chain configuration, and turns into gel that looks white when boiled up to 70 degrees C. He explains amino acids building egg-white protein are connected with -CO-NH- segments upon which a large number of water molecules are attached by a linear association, as such CO\*OH<sub>2</sub>\*H<sub>2</sub>O\*OH<sub>2</sub>\*H<sub>2</sub>O\*....., and NH\*H<sub>2</sub>O\*OH<sub>2</sub>\*H<sub>2</sub>O\*..... He defines this water chains assciated linearly by Meissoner's magnetism as bio-water. So that protein chains with such a long dimension can be coexist with water molecules, or can dissolve in water, as though caffeine and nicotine molecules can dissolve in CO<sub>2</sub> gas phase to form a gas/solid phase.

Biochemistry today explaining this dissolution of egg-white into water, attributes to protein chains with globular folding. However, it is wrong. When it is heated more than 70 degrees C, or mechanically agitated these bio-water chains are stripped off, and the bare protein chains that are inherently hydrophobic, have to precipitate and aggregate by themselves building a network structure to give a gel phase. Thus liberated water molecules from raw egg-white become to pack in the protein network holes as tiny droplets, scattering the incoming light to be seen white. We can observe simply this mysterious phenomenon if we dilute this egg-white in water and heat it slowly up to 70 degree C, or agitate it vigorously. However, the polymer chains from which water chains are stripped off, never redissolve into their initial water solvent. it is because the space between aggregated protein chains does not allow to rebuild bio-water chains, due to stronger attraction between them. Likewise silk fibrin protein attaching bio-waters on their CO-NH segments are contained in silk worm body. When this aqueous silk fibrin is discharged as filaments, it builds silk fibers evaporating water, but they never redissolve into the same water solvent. Thus this kind of worms can build their own cocoons possible to protect their pupas from rains and external enemy

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attack. newedama wrote

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• **Follow-Ups:**

- ◆ **Re: How can NaCl dissolve in water? Can Any one explain this phenomenon in molecular level?**

◇ From: Dirk Van de moortel

• **References:**

- ◆ **How can NaCl dissolve in water? Can Any one explain this phenomenon in molecular level?**

◇ From: newedana

- ◆ **Re: How can NaCl dissolve in water? Can Any one explain this phenomenon in molecular level?**

◇ From: pete

• Prev by Date: **Re: Yet Another Rail Gun Question**

• Next by Date: **Re: John Polasek "Dual Space" (real or imaginary??)**

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